

SECTION C — CHEMISTRY; METALLURGY

C01 INORGANIC CHEMISTRY

C01F COMPOUNDS OF THE METALS BERYLLIUM, MAGNESIUM, ALUMINIUM, CALCIUM, STRONTIUM, BARIUM, RADIUM, THORIUM, OR OF THE RARE-EARTH METALS (metal hydrides C01B 6/00; salts of oxyacids of halogens C01B 11/00; peroxides, salts of peroxyacids C01B 15/00; sulfides or polysulfides of magnesium, calcium, strontium, or barium C01B 17/42; thiosulfates, dithionites, polythionates C01B 17/64; compounds containing selenium or tellurium C01B 19/00; binary compounds of nitrogen with metals C01B 21/06; azides C01B 21/08; metal amides C01B 21/092; nitrites C01B 21/50; phosphides C01B 25/08; salts of oxyacids of phosphorus C01B 25/16; carbides C01B 31/30; compounds containing silicon C01B 33/00; compounds containing boron C01B 35/00; compounds having molecular sieve properties but not having base-exchange properties C01B 37/00; compounds having molecular sieve and base-exchange properties, e.g. crystalline zeolites, C01B 39/00; cyanides C01C 3/08; salts of cyanic acid C01C 3/14; salts of cyanamide C01C 3/16; thiocyanates C01C 3/20; fermentation or enzyme-using processes for the preparation of elements or inorganic compounds except carbon dioxide C12P 3/00; obtaining metal compounds from mixtures, e.g. ores, which are intermediate compounds in a metallurgical process for obtaining a free metal C22B; production of non-metallic elements or inorganic compounds by electrolysis or electrophoresis C25B)

Note(s)

- Attention is drawn to Note (1) after class C01, which defines the last place priority rule applied in this class, i.e. in the range of subclasses C01B-C01G and within these subclasses.
- Therapeutic activity of compounds is further classified in subclass A61P.

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| 1/00 | Methods of preparing compounds of the metals beryllium, magnesium, aluminium, calcium, strontium, barium, radium, thorium, or the rare earths, in general | 5/38 | • Magnesium nitrates |
| | | 5/40 | • Magnesium sulfates (double sulfates of magnesium with sodium or potassium C01D 5/12, with other alkali metals C01D 15/06, C01D 17/00) [3] |
| | | 5/42 | • Magnesium sulfites |
| 3/00 | Compounds of beryllium | 7/00 | Compounds of aluminium |
| 3/02 | • Oxides; Hydroxides [3] | 7/02 | • Aluminium oxide; Aluminium hydroxide; Aluminates |
| 5/00 | Compounds of magnesium | 7/04 | • • Preparation of alkali metal aluminates; Aluminium oxide or hydroxide therefrom |
| 5/02 | • Magnesia | 7/06 | • • • by treating aluminous minerals with alkali hydroxide |
| 5/04 | • • by oxidation of metallic magnesium | 7/08 | • • • by treating aluminous minerals with sodium carbonate |
| 5/06 | • • by thermal decomposition of magnesium compounds (calcining magnesite or dolomite C04B 2/10) | 7/10 | • • • by treating aluminous minerals with alkali sulfates and reducing agents |
| 5/08 | • • • by calcining magnesium hydroxide | 7/12 | • • • Alkali metal aluminates from alkaline earth metal aluminates |
| 5/10 | • • • by thermal decomposition of magnesium chloride with water vapour | 7/14 | • • • Aluminium oxide or hydroxide from alkali metal aluminates |
| 5/12 | • • • by thermal decomposition of magnesium sulfate, with or without reduction | 7/16 | • • Preparation of alkaline earth metal aluminates; Aluminium oxide or hydroxide therefrom |
| 5/14 | • Magnesium hydroxide | 7/18 | • • • Aluminium oxide or hydroxide from alkaline earth metal aluminates |
| 5/16 | • • by treating magnesia, e.g. calcined dolomite, with water or solutions of salts not containing magnesium | 7/20 | • • Preparation of aluminium oxide or hydroxide from aluminous ores with acids or salts |
| 5/20 | • • by precipitation from solutions of magnesium salts with ammonia | 7/22 | • • • with halides |
| 5/22 | • • from magnesium compounds with alkali hydroxides or alkaline earth oxides or hydroxides | 7/24 | • • • with nitric acid or nitrogen oxides |
| 5/24 | • Magnesium carbonates | 7/26 | • • • with sulfuric acids or sulfates |
| 5/26 | • Magnesium halides | 7/28 | • • • with sulfurous acid |
| 5/28 | • • Fluorides | 7/30 | • • Preparation of aluminium oxide or hydroxide by thermal decomposition of aluminium compounds |
| 5/30 | • • Chlorides | 7/32 | • • • of sulfates |
| 5/32 | • • • Preparation of anhydrous magnesium chloride by chlorinating magnesium compounds | 7/34 | • • Preparation of aluminium hydroxide by precipitation from solutions containing aluminium salts |
| 5/34 | • • • Dehydrating magnesium chloride containing water of crystallisation | | |
| 5/36 | • • Bromides | | |

C01F

- 7/36 • • • from organic aluminium salts
- 7/38 • • Preparation of aluminium oxide by thermal reduction of aluminous minerals
- 7/40 • • • in the presence of aluminium sulfide
- 7/42 • • Preparation of aluminium oxide or hydroxide from metallic aluminium, e.g. by oxidation
- 7/44 • • Dehydration of aluminium hydroxide
- 7/46 • • Purification of aluminium oxide, aluminium hydroxide or aluminates [5]
- 7/47 • • • of aluminates [5]
- 7/48 • Aluminium halides
- 7/50 • • Fluorides
- 7/52 • • • Double compounds containing both fluorine and other acid groups
- 7/54 • • • Double compounds containing both aluminium and alkali metals or alkaline earth metals
- 7/56 • • Chlorides (containing fluorine C01F 7/52) [3]
- 7/58 • • • Preparation of anhydrous aluminium chloride
- 7/60 • • • • from oxygen-containing aluminium compounds
- 7/62 • • • Purification
- 7/64 • • Bromides (containing fluorine C01F 7/52) [3]
- 7/66 • Aluminium nitrates (containing fluorine C01F 7/52) [3]
- 7/68 • Aluminium compounds containing sulfur (containing fluorine C01F 7/52) [3]
- 7/70 • • Sulfides
- 7/72 • • Sulfites
- 7/74 • • Sulfates
- 7/76 • • • Double salts, e.g. alums
- 11/00 Compounds of calcium, strontium, or barium** (C01F 7/00 takes precedence) [3]
- 11/02 • Oxides or hydroxides (production of lime C04B 2/00)
- 11/04 • • by thermal decomposition
- 11/06 • • • of carbonates
- 11/08 • • by reduction of sulfates
- 11/10 • • from sulfides
- 11/12 • • from silicates
- 11/16 • • Purification
- 11/18 • Carbonates
- 11/20 • Halides
- 11/22 • • Fluorides
- 11/24 • • Chlorides
- 11/26 • • • from sulfides
- 11/28 • • • by chlorination of alkaline earth metal compounds
- 11/30 • • • Concentrating; Dehydrating; Preventing the absorption of moisture or caking
- 11/32 • • • Purification
- 11/34 • • Bromides
- 11/36 • Nitrates
- 11/38 • • Preparation with nitric acid or nitrogen oxides
- 11/40 • • Preparation by double decomposition with nitrates
- 11/42 • • Double salts (with magnesium C01F 5/38)
- 11/44 • • Concentrating; Crystallising; Dehydrating; Preventing the absorption of moisture or caking
- 11/46 • Sulfates (dehydration of gypsum C04B 11/02)
- 11/48 • Sulfites
- 13/00 Compounds of radium**
- 15/00 Compounds of thorium**
- 17/00 Compounds of the rare-earth metals, i.e. scandium, yttrium, lanthanum, or the group of the lanthanides**